

Multiple Choice: Hybridization

1. What is the hybridization of the central atom in XeF_4 ?

- A) sp^2 B) sp^3 C) sp^3d D) sp^3d^2

ANS: D

1. Carbon uses _____ hybrid orbitals in CICH_3 .

- A) sp B) sp^2 C) sp^3 D) sp^3d

ANS: A

2. Valence bond theory predicts that tin will use _____ hybrid orbitals in SnCl_4 .

- A) sp B) sp^2 C) sp^3 D) sp^3d

ANS: C

3. Valence bond theory predicts that xenon will use _____ hybrid orbitals in XeOF_4 .

- A) sp^3 B) sp^3 C) sp^3d D) sp^3d^2

ANS: D

31. What is the hybridization in a molecule with 120° bond angles exclusively?

- A) sp B) sp^2 C) sp^3 D) sp^3d E) sp^3d^2

ANS: B

33. Which hybridization scheme results in a planar molecule (ie, all atoms lying in the same plane if there are no unshared electron pairs around the central atom)?

- A) sp^2 B) sp^3 C) sp^3 D) sp^3d E) sp^3d^2

ANS: A

26. Which of the underlined atoms (C_1 , C_2 , N , and O) are sp^2 hybridized?



- A) C_1 and C_2 only B) C_2 , N , and O C) N and O D) O and C_2 E) O

12 Hsc Chemistry Multiple Choice Answers

J. Compton, Allyn Jones



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