

QUESTION 2

Vitash works at a chemical company and has been asked to design an industrial process to produce sulfuric acid.

- 2.1 One of the reactions in the production of sulfuric acid is the roasting (heating in oxygen) of a metal ore that contains lead(II) sulfide:



Vitash does a test experiment in which 38,8 g of O_2 gas completely reacts with 800 g of the metal ore. All of the PbS in the ore reacts, and ONLY the PbS in the ore reacts with the oxygen.

- 2.1.1 Define *molar mass*. (2)
- 2.1.2 Calculate the amount (in moles) of O_2 that reacted. (3)
- 2.1.3 Calculate the mass of pure PbS in the metal ore. (3)
- 2.1.4 Hence, calculate the mass percentage of the PbS in the metal ore. (2)

- 2.2 In another test experiment, Vitash reacts 50 g of sulfur trioxide with water.

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Richard Bailey



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