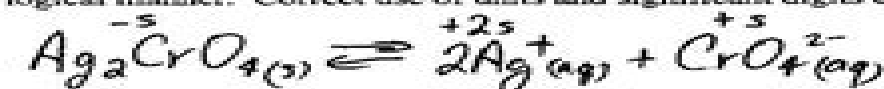


10. Calculate the molar solubility of silver chromate in water at 25°C. Show all your steps in a logical manner. Correct use of units and significant digits counts (3 marks)



$$K_{sp} = [\text{Ag}^+]^2 [\text{CrO}_4^{2-}]$$

$$K_{sp} = (2s)^2 s$$

$$K_{sp} = 4s^3$$

$$s^3 = \frac{K_{sp}}{4} \quad \text{or} \quad s = \sqrt[3]{\frac{K_{sp}}{4}}$$

$$= \sqrt[3]{\frac{1.1 \times 10^{-12}}{4}}$$

$$= 6.5 \times 10^{-5} \text{ M}$$

Answer $6.5 \times 10^{-5} \text{ M}$

11. At a certain temperature 0.0558 grams of SrF_2 will dissolve in 500.0 mL of water. Calculate the K_{sp} for SrF_2 at this temperature. Show all your steps in a logical manner. Correct use of units and significant digits counts (3 marks)

$$\text{g} \rightarrow \text{mol} \rightarrow \text{M}$$

$$0.0558 \text{ g SrF}_2 \times \frac{1 \text{ mol}}{125.6 \text{ g}} = 4.4427 \times 10^{-4} \text{ mol}$$

$$\text{M (molar solubility)} = \frac{4.4427 \times 10^{-4} \text{ mol}}{0.5000 \text{ L}} = 8.885 \times 10^{-4} \text{ M}$$



$$K_{sp} = [\text{Sr}^{2+}][\text{F}^-]^2$$

$$= (8.885 \times 10^{-4})(1.777 \times 10^{-3})^2$$

Answer $K_{sp} = 2.81 \times 10^{-9}$

12. Which is most soluble in water at 25 °C, lead (II) bromide, lead (II) chloride, lead (II) iodide, or lead (II) iodate? (2 marks)

Answer PbCl_2 . How did you obtain your answer? All the same type (AB_2) PbCl_2 has the highest K_{sp}

13. Which is least soluble in water at 25 °C, lead (II) bromide, lead (II) chloride, lead (II) iodide, or lead (II) iodate? (2 marks)

Answer $\text{Pb}(\text{IO}_3)_2$. How did you obtain your answer? All the same type. $\text{Pb}(\text{IO}_3)_2$ has the lowest K_{sp} .

10
10

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Ashok Kumar Jha



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Calculating Theoretical Yield Is Not Impossible! Michael S. Leonard, 2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory. This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill. It begins with simple examples from everyday life, demonstrates the importance of balancing the equation, addresses the role of the mole in these computations, discusses different types of liquids, considers the role of significant figures, and culminates with the planning of syntheses. There are suggestions for further reading as well as practice problems and questions to ensure mastery. Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory. Addresses the necessity of balancing the reaction equation, the centrality of the mole in these calculations, and the role of significant figures in reporting the answer. Explains how to approach the calculations when using neat liquids or solutions. The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product. All of the problems in the book include detailed solutions with accompanying text to explain the answers, and ancillaries also include suggestions for further reading.

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