

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization constant | 5. K_a | 6. K_b |
| | 7. weak solution | |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constant(s) that are indicated in the spaces provided. Be sure to include correct units when appropriate.

1. Water autoionizes. K_w is the ionization constant of water. Write the equation and K_w at 25°C.
2. The autoionization constant of water is 1.0×10^{-14} at 25°C. Write the equation and K_w at 25°C.
3. The autoionization constant of water is 1.0×10^{-14} at 25°C. Write the equation and K_w at 25°C.
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9. The autoionization constant of water is 1.0×10^{-14} at 25°C. Write the equation and K_w at 25°C.
10. The autoionization constant of water is 1.0×10^{-14} at 25°C. Write the equation and K_w at 25°C.

1. $K_w = 1.0 \times 10^{-14}$ 2. $K_w = 1.0 \times 10^{-14}$ 3. $K_w = 1.0 \times 10^{-14}$ 4. $K_w = 1.0 \times 10^{-14}$ 5. $K_w = 1.0 \times 10^{-14}$ 6. $K_w = 1.0 \times 10^{-14}$ 7. $K_w = 1.0 \times 10^{-14}$ 8. $K_w = 1.0 \times 10^{-14}$ 9. $K_w = 1.0 \times 10^{-14}$ 10. $K_w = 1.0 \times 10^{-14}$ 11. $K_w = 1.0 \times 10^{-14}$ 12. $K_w = 1.0 \times 10^{-14}$

19 2 Hydrogen Ions And Acidity Section Review Answers

Kevin T. Patton, Gary A. Thibodeau



19 2 Hydrogen Ions And Acidity Section Review Answers:

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