

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. acid rain | 6. acid rain |
| 7. autoionization | 8. acid rain | 9. pH |

Key Equation

- | | |
|------------------------------------|------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}^+]$ |

Part A Completion

Write the equation for the autoionization of water. What is the pH of pure water at 25°C? What is the pOH of pure water at 25°C? What is the pH of pure water at 25°C?

1. What is the pH of a 0.1 M solution of hydrochloric acid?
2. What is the pH of a 0.1 M solution of sodium hydroxide?
3. What is the pH of a 0.1 M solution of acetic acid?
4. What is the pH of a 0.1 M solution of ammonia?
5. What is the pH of a 0.1 M solution of hydrofluoric acid?
6. What is the pH of a 0.1 M solution of hydrocyanic acid?
7. What is the pH of a 0.1 M solution of hydrobromic acid?
8. What is the pH of a 0.1 M solution of hydroiodic acid?
9. What is the pH of a 0.1 M solution of hydrochloric acid?
10. What is the pH of a 0.1 M solution of hydrobromic acid?
11. What is the pH of a 0.1 M solution of hydroiodic acid?
12. What is the pH of a 0.1 M solution of hydrobromic acid?
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40. What is the pH of a 0.1 M solution of hydrobromic acid?

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Michael Rajnik



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Table of Contents 19 2 Hydrogen Ions And Acidity Section Review Answers

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Section Review Answers
 - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Hydrogen Ions And Acidity Section Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Personalized Recommendations
 - 19 2 Hydrogen Ions And Acidity Section Review Answers User Reviews and Ratings
 - 19 2 Hydrogen Ions And Acidity Section Review Answers and Bestseller Lists
5. Accessing 19 2 Hydrogen Ions And Acidity Section Review Answers Free and Paid eBooks

- 19 2 Hydrogen Ions And Acidity Section Review Answers Public Domain eBooks
 - 19 2 Hydrogen Ions And Acidity Section Review Answers eBook Subscription Services
 - 19 2 Hydrogen Ions And Acidity Section Review Answers Budget-Friendly Options
6. Navigating 19 2 Hydrogen Ions And Acidity Section Review Answers eBook Formats
- ePub, PDF, MOBI, and More
 - 19 2 Hydrogen Ions And Acidity Section Review Answers Compatibility with Devices
 - 19 2 Hydrogen Ions And Acidity Section Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Interactive Elements 19 2 Hydrogen Ions And Acidity Section Review Answers
8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Section Review Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Section Review Answers
9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Section Review Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Section Review Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Section Review Answers
- Setting Reading Goals 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Section Review Answers
- Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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