

19.5

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization constant | 5. acid rain | 6. neutralization |
| | 7. buffer solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|--|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pOH} = -\log [\text{OH}^-]$ |
| 3. $\text{pH} + \text{pOH} = 14$ | 4. $\text{pH} = \text{pK}_a + \log \frac{[\text{A}^-]}{[\text{HA}]}$ |

Part A Completion

Write the requested unit or numerical information of the values provided that are indicated in parentheses. Significant figures represent values that are completely arbitrary.

1. What substance is _____, a weak hydrogen ion donor?
2. Hydrochloric acid (HCl) is the conjugate acid of what which provides an effective buffer system? _____
3. Hydrochloric acid ionizes from _____ to a 4.0% extent.
4. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
5. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
6. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
7. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
8. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
9. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
10. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
11. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
12. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
13. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
14. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
15. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
16. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
17. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
18. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
19. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.
20. _____ ionizes to a 4.0% extent with the acid ionization constant $K_a = 1.0 \times 10^{-4}$.

1. _____

2. _____

3. _____

4. _____

5. _____

6. _____

7. _____

8. _____

9. _____

10. _____

11. _____

12. _____

13. _____

14. _____

15. _____

16. _____

17. _____

18. _____

19. _____

20. _____

19 2 Hydrogen Ions And Acidity Work Sheet Answers

Michael Brown



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