

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain |
| 7. autoionization | 8. acid rain | 9. acid |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the unknown quantity that are indicated in the problem. Be sure to include units whenever a unit is required.

1. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.
2. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.
3. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.
4. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.
5. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.
6. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.
7. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.

8. _____
9. _____
10. _____
11. _____
12. _____
13. _____
14. _____
15. _____
16. _____
17. _____
18. _____

19 2 Hydrogen Ions And Acidity Answer Key

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