

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on  $K_a$  and  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                    |
|-------------------|-----------------------------|--------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base  |
| 4. autoionization | 5. acid rain                | 6. buffer solution |
| 7. ionization     | 8. weak solution            | 9. pH              |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) for the reaction(s) described. Show the calculation for pH or percent ionization when required. Assume acids are monoprotic unless noted.

1. What substance is \_\_\_\_\_, a weak hydrogen ion donor?
2. \_\_\_\_\_ is the concentration of those species present in solution that equal to \_\_\_\_\_.
3. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
4. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
5. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
6. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
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11. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
12. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
13. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
14. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
15. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
16. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
17. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
18. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
19. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
20. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.

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# 19 2 Hydrogen Ions And Acidity Answer Key

**David Baud**



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