

SCIENCE OF COMBUSTION

While dealing with combustion examples, we have already talked about the combustion with scientific outlook. Let us recap it and explore it little further.

Combustion is a **chemical reaction** between a *fuel* and an *oxidizing agent*, typically *oxygen*, that produces **heat**, **light**, and often other byproducts such as **water vapor** and **carbon dioxide**. It is a fundamental process that occurs in many natural and human-made systems, ranging from the *burning of wood* in a campfire to the *combustion of gasoline* in an automobile engine.

According to the National Institute of Standards and Technology (NIST), combustion is a rapid oxidation process that releases heat and light, and is often accompanied by a change in color and the formation of smoke.

The following are the three main descriptions of combustion, as per the NIST definition:

1. A combustion reaction is a process that is exothermic, irreversible, and produces heat and light.
2. The rate of combustion defines combustion, because, if the reaction is fast, it is called combustion, and if it is slow, it is called oxidation.
3. There is a minimum amount of fuel required for combustion, known as the lower flammable limit (LFL).
4. Light is an important product of the combustion, which is often visible as a flame or a fire.
5. Heat is an important product of the combustion, which is often visible as a flame or a fire.

Key characteristics of combustion

There are three main characteristics of combustion:

Exothermic Reaction

Combustion is an exothermic reaction, meaning it releases heat. The heat released is often used to power engines or to heat buildings. The general chemical equation for the combustion of hydrocarbons is as follows:

$$\text{Fuel} + \text{Oxygen} \rightarrow \text{Products} + \text{Heat}$$

For example,



Heat Release

Combustion is an exothermic reaction, meaning it releases energy. This energy is responsible for the heat of the flame and is responsible for raising the temperature of the surrounding environment.

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S Ben Porath



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