

Acid-Base Titration Lab

Introduction

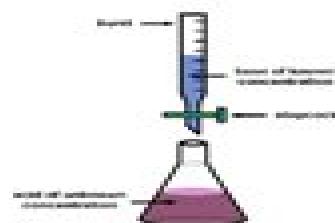
In chemistry laboratory, it is sometimes necessary to experimentally determine the concentration of an unknown acid or base solution. A procedure for making this kind of determination is called an **acid-base titration**. In this laboratory process, a solution of known concentration, called the standard solution, is carefully added to a solution of unknown concentration until the mixture becomes neutral. The neutral point of the solution is recognized by an indicator's color change. If the unknown solution is acidic, then the standard solution will be basic. The opposite would be true if the unknown solution was basic.

We know that the mixing of equal amounts of acid and base ions will create neutral water. At the molecular level, this reaction can be illustrated with the following equation.



This equation states that one mole of hydrogen ions (acid) will neutralize one mole of hydroxide ions (base). Since we can exactly measure the moles of the standard solution, we can assume that the moles of the solution of unknown concentration will be the same at the neutral point. This is called the end-point of the titration. Using the equation $M_aV_a = M_bV_b$, we can use the experimental data from the titration to find the unknown concentration.

Objective: In this experiment the unknown solution will be HCl(aq) and the standard solution will be the base sodium hydroxide. You will know the concentration of the base and the volume of the acid and base used. With this information you can use the titration formula to calculate the concentration of the acid. The diagram below shows the set up.



Move your eyes to the level of the liquid surface and read the value of the bottom of the meniscus with 1/10 of the smallest scale marked on the buret.

Swirl the flask continuously until one drop of titrant causes a color change to remain.
A white sheet of paper can be placed under the flask to help you recognize the color change at the end-point.

Pre Lab

1. Define the following words:
Titration- _____
Endpoint- _____
Neutralization- _____
2. Write the neutralization reaction for HCl reacting with NaOH.
3. What is the pH of the solution at the end point of the titration?

Procedure

1. The flask is filled with 10 mL of unknown concentration of HCl. **Record the volume of acid on your data form**
2. Phenolphthalein Indicator is added to the flask.
3. **Record on your data form** the initial volume of base in the buret
4. You will now start adding base (0.25M NaOH) from the buret into the flask to neutralize the acid. The flask is gently swirled as the base is added. A pink color should appear as the base is added, but it will disappear as the flask is swirled. The end-point of the titration will be reached when 1 drop of base makes the solution turn and stay pink.
5. **Record the final volume of base** from the buret that was needed to neutralize the acid.

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Tao Wei



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Table of Contents Acid Base Titration Lab Chem Fax Answers

1. Understanding the eBook Acid Base Titration Lab Chem Fax Answers
 - The Rise of Digital Reading Acid Base Titration Lab Chem Fax Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Titration Lab Chem Fax Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid Base Titration Lab Chem Fax Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Base Titration Lab Chem Fax Answers
 - Personalized Recommendations
 - Acid Base Titration Lab Chem Fax Answers User Reviews and Ratings
 - Acid Base Titration Lab Chem Fax Answers and Bestseller Lists
5. Accessing Acid Base Titration Lab Chem Fax Answers Free and Paid eBooks
 - Acid Base Titration Lab Chem Fax Answers Public Domain eBooks
 - Acid Base Titration Lab Chem Fax Answers eBook Subscription Services
 - Acid Base Titration Lab Chem Fax Answers Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Acid Base Titration Lab Chem Fax Answers Compatibility with Devices
 - Acid Base Titration Lab Chem Fax Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Acid Base Titration Lab Chem Fax Answers
 - Interactive Elements Acid Base Titration Lab Chem Fax Answers
8. Staying Engaged with Acid Base Titration Lab Chem Fax Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid Base Titration Lab Chem Fax Answers
9. Balancing eBooks and Physical Books Acid Base Titration Lab Chem Fax Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid Base Titration Lab Chem Fax Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Acid Base Titration Lab Chem Fax Answers
 - Setting Reading Goals Acid Base Titration Lab Chem Fax Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acid Base Titration Lab Chem Fax Answers
 - Fact-Checking eBook Content of Acid Base Titration Lab Chem Fax Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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his son jan the younger continued the partnership with rubens until the latter's death in 1640

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