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Chapter 2

The First Law of Thermodynamics

- Work is defined as negative pΔV because if a system does work on the surroundings, the system loses energy.
- 2.4. $1780 \text{torr} \left| \frac{1 \text{atm}}{} \right| = 2.34 \text{atm}$ $\left(760 \text{torr} \right)$ $w = -P \Delta V = -2.34 \text{atm} \left(1.00 L 3.55 L \right) \left(\frac{101.32 J}{1 L_{\odot} \text{atm}} \right) = +605 J$

605 J of work are done ON the system.

- (a) The work would be less because the external pressure is less.
 - (b) The work would be greater because the external pressure is greater.
 - (c) No work would be performed because the external pressure is (effectively) zero.
- First, we need to find the final volume of the CO₂. Using P₁V₁=P₂V₂, we find that V₂=105mL.

$$w = -(1.0atm)(105 - 25.0mL)\left(\frac{1L}{1000mL}\right)\left(\frac{101.32J}{1L \cdot atm}\right) = -8.1J$$

2.10. First, determine the ∆T: 330 K − 298 K = 32 K.

Using
$$q = m \cdot c \cdot \Delta T$$
, we rearrange: $c = \frac{q}{m \cdot \Delta T}$

Substituting:
$$c = \frac{288 \text{ J}}{(50.5 \text{ g})(32 \text{ K})} = 0.178 \frac{\text{J}}{\text{g} \cdot \text{K}}$$
.

2.12. First, we need to calculate the number of moles of phosphorous:

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