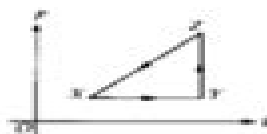


**Sample Multiple Choice Questions from past AP Physics B tests –
Atomic, Molecular, Thermal**

1984 B multiple choice

30. A 2-kilogram block of metal with a specific heat of 100 joules per kilogram*kelvin falls from rest through a distance of 100 meters to the Earth's surface. If half of the potential energy lost by the fallen block is converted to internal energy of the block, the temperature change of the block is most nearly
(A) 1 K (B) 5 K (C) 10 K (D) 25 K (E) 45 K
31. The maximum efficiency of a heat engine that operates between temperatures of 1500 K in the firing chamber and 600 K in the exhaust chamber is most nearly (A) 33% (B) 40% (C) 60% (D) 67% (E) 100%
32. An ideal gas is made up of N diatomic molecules, each of mass M . All of the following statements about this gas are true EXCEPT:
(A) The temperature of the gas is proportional to the average translational kinetic energy of the molecules.
(B) All of the molecules have the same speed.
(C) The molecules make elastic collisions with the walls of the container.
(D) The molecules make elastic collisions with each other.
(E) The average number of collisions per unit time that the molecules make with the walls of the container depends on the temperature of the gas.



A thermodynamic system is taken from an initial state X along the path XYZ as shown in the PV -diagram above.

33. For the process $X \rightarrow Y$, ΔU is greater than zero and
(A) $Q < 0$ and $W = 0$ (B) $Q < 0$ and $W > 0$ (C) $Q > 0$ and $W < 0$ (D) $Q > 0$ and $W > 0$ (E) $Q < 0$ and $W > 0$
34. For the process $Y \rightarrow Z$, Q is greater than zero and
(A) $W < 0$ and $\Delta U = 0$ (B) $W < 0$ and $\Delta U < 0$ (C) $W > 0$ and $\Delta U > 0$ (D) $W > 0$ and $\Delta U = 0$ (E) $W > 0$ and $\Delta U < 0$
50. An ideal gas confined in a box initially has pressure p . If the absolute temperature of the gas is doubled and the volume of the box is quadrupled, the pressure is (A) $p/8$ (B) $p/4$ (C) $p/2$ (D) p (E) $2p$



68. Heat is added at a constant rate to a sample of pure substance that is initially a solid at temperature T_1 . The temperature of the sample as a function of time is shown in the graph above. From the graph, one can conclude that the
(A) substance sublimates directly from the solid phase to the vapor phase

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