

LIFE SCIENCES - Grade 10**TOPIC TEST: History of Life on Earth****MARKS:30****TIME: 30 minutes**

SECTION A**QUESTION 1**

Various options are given as possible answers to the following questions. Choose the answer and write only the letter (**A to D**) next to the question number (**1.1 to 1.5**) in the **ANSWER BOOK**, for example **1.6 D**.

- 1.1 Palaeontology is a study of ...
- A the half-life of radioactive substances.
 - B continental drift.
 - C fossils.
 - D changes during the Cambrian explosion.
- 1.2 Which method allows scientists to say that one fossil is older than the other?
- A Fossilization
 - B Relative dating
 - C Petrification
 - D Radiometric dating
- 1.3 A period in the earth's history where global temperatures plummet and there is extensive glaciation is known as ...
- A the greenhouse effect.
 - B a cold spell.
 - C a hole in the ozone.
 - D an ice age.
- 1.4are the richest sources of fossils
- A Layers of strata
 - B Index fossils
 - C Mineral salts
 - D Sedimentary rocks

13 Grade 10 Life Sciences Question Paper

TD Snyder



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web solubility data task complete the model use google spreadsheets to make a graph of the solubility data in table create a google sheet copy and paste the data table above into sheets

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web pogil solubility why solubility is a measure of the amount of solute that will dissolve in a given amount of solvent usually water at a given temperature a solubility curve shows how much solute dissolves in a given volume of solvent at a given temperature

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web information solubility rules a solute is considered soluble if an appreciable amount of it can be dissolved in a given amount of the solvent for example both table salt nacl and table sugar c₁₂h₂₂o₁₁ are soluble substances in water a solute is considered insoluble if very little of it dissolves in a given amount of the solvent

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web predict what might happen to the solubility of the insoluble salt if a 1.0 M acid solution was added drop wise to the beaker hint which ion species in the beaker will react with the acid how will this reaction alter the concentration of the common ion