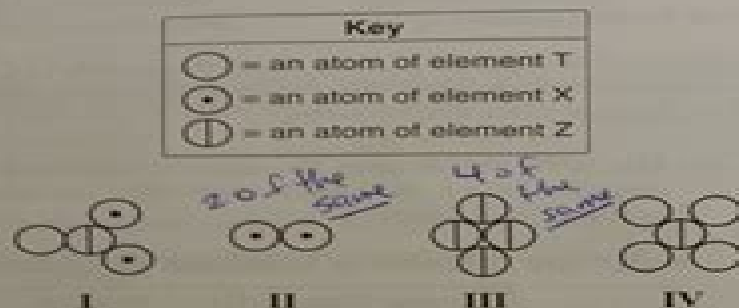


39 Given four particle models:



Which two models can be classified as elements?

- (1) I and II (3) II and III
(2) I and IV (4) II and IV

40 After being thoroughly stirred at 10.°C, which mixture is heterogenous?

- (1) 25.0 g of KCl and 100. g of H₂O *unsat.*
(2) 25.0 g of KNO₃ and 100. g of H₂O *supersat.*
(3) 25.0 g of NaCl and 100. g of H₂O *unsat.*
(4) 25.0 g of NaNO₃ and 100. g of H₂O *unsat.*

41 Which two compounds are electrolytes?

- (1) KOH and CH₃COOH
(2) KOH and C₃H₁₂
(3) CH₃OH and CH₃COOH
(4) CH₃OH and C₃H₁₂

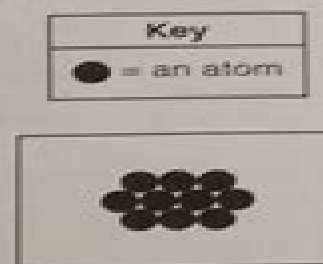
42 Which statement explains why a CO₂ molecule is nonpolar?

- (1) Carbon and oxygen are both nonmetals.
(2) Carbon and oxygen have different electronegativities.
(3) The molecule has a symmetrical distribution of charge.
(4) The molecule has an asymmetrical distribution of charge.

43 Which temperature change indicates an increase in the average kinetic energy of the molecules in a sample?

- (1) 15°C to 298 K ↑ (3) 305 K to 0°C ↓
(2) 37°C to 273 K ↓ (4) 355 K to 25°C ↓

44 Given the particle diagram:



Which substance at STP can be represented by this particle diagram? *Solid*

- (1) N₂ (O₂) (3) Mg (Z)
(2) H₂ (O₂) (4) Kr (O₂)

45 Which type of equilibrium exists in a sealed flask containing Br₂(l) and Br₂(g) at 298 K and 1.0 atm?

- (1) static phase equilibrium
(2) static solution equilibrium
(3) dynamic phase equilibrium
(4) dynamic solution equilibrium

46 What are the products when potassium hydroxide reacts with hydrochloric acid?

- (1) KH(s), Cl⁺(aq), and OH⁻(aq)
(2) K(s), Cl₂(g), and H₂O(l)
(3) KCl(aq) and H₂O(l)
(4) KOH(aq) and Cl₂(g)

47 In a titration, 20.0 milliliters of a 0.150 M NaOH(aq) solution exactly neutralizes 24.0 milliliters of an HCl(aq) solution. What is the concentration of the HCl(aq) solution?

- (1) 0.125 M (3) 0.250 M
(2) 0.180 M (4) 0.360 M

48 What fraction of a Sr-90 sample remains unchanged after 87.3 years?

- (1) $\frac{1}{2}$ (3) $\frac{1}{4}$
(2) $\frac{1}{3}$ (4) $\frac{1}{8}$

Handwritten calculation:

$$3 = \frac{87.3}{27.14}$$

$$\frac{1}{2} = \frac{1}{8}$$

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Xiaolong Qi



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