

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and concentrations of acids and bases and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain |
| 7. autoionization | 8. acid rain | 9. pH |

Key Equation

- | | |
|------------------------------------|------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}^+]$ |

Part A Completion

Write the equation and/or the chemical formula(s) of the species involved in the reaction described in each problem. Be sure to include the correct state symbols.

1. Write the equation for the autoionization of water.
2. Write the equation for the autoionization of water.
3. Write the equation for the autoionization of water.
4. Write the equation for the autoionization of water.
5. Write the equation for the autoionization of water.
6. Write the equation for the autoionization of water.
7. Write the equation for the autoionization of water.
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9. Write the equation for the autoionization of water.
10. Write the equation for the autoionization of water.
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12. Write the equation for the autoionization of water.
13. Write the equation for the autoionization of water.
14. Write the equation for the autoionization of water.
15. Write the equation for the autoionization of water.
16. Write the equation for the autoionization of water.
17. Write the equation for the autoionization of water.
18. Write the equation for the autoionization of water.
19. Write the equation for the autoionization of water.
20. Write the equation for the autoionization of water.

1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
4. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
5. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
6. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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8. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
9. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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11. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
12. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
13. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
14. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
15. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
16. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
17. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
18. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
19. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
20. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

19 2 Hydrogen Ions And Acidity Work Sheet Answers

Ian Pickup



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web i don t think the new csm codex will let you take rubrics in a black legion army but it might let you bring them as allies or maybe in a specialist detachment but if you really want to

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designed work alongside the rules found in codex

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