

ACID-BASE RELATIVE STRENGTHS

in Aqueous Solutions at 25° C

Name of Acid	Acid		Base		K_a
Hydriodic	HI	→	$H^+ + I^-$		$\gg 1$
Hydrobromic	HBr	→	$H^+ + Br^-$		$\gg 1$
Hydrochloric	HCl	→	$H^+ + Cl^-$		$\gg 1$
Nitric	HNO_3	→	$H^+ + NO_3^-$		> 1
Sulfuric	H_2SO_4	→	$H^+ + HSO_4^-$		> 1
Oxalic	$H_2C_2O_4$	$\rightleftharpoons$	$H^+ + HC_2O_4^-$		5.90×10^{-2}
Sulfurous	H_2SO_3	$\rightleftharpoons$	$H^+ + HSO_3^-$		1.54×10^{-2}
Hydrogen Sulfate Ion	HSO_4^-	$\rightleftharpoons$	$H^+ + SO_4^{2-}$		1.20×10^{-2}
Phosphoric	H_3PO_4	$\rightleftharpoons$	$H^+ + H_2PO_4^-$		7.52×10^{-3}
Nitrous	HNO_2	$\rightleftharpoons$	$H^+ + NO_2^-$		4.6×10^{-4}
Hydrofluoric	HF	$\rightleftharpoons$	$H^+ + F^-$		3.53×10^{-4}
Formic	HCO_2H	$\rightleftharpoons$	$H^+ + HCO_2^-$		1.77×10^{-4}
Benzoic	$C_6H_5CO_2H$	$\rightleftharpoons$	$H^+ + C_6H_5CO_2^-$		6.46×10^{-5}
Acetic	CH_3CO_2H	$\rightleftharpoons$	$H^+ + CH_3CO_2^-$		1.76×10^{-5}
Carbonic	H_2CO_3	$\rightleftharpoons$	$H^+ + HCO_3^-$		4.30×10^{-7}
Hydrogen Sulfite Ion	HSO_3^-	$\rightleftharpoons$	$H^+ + SO_3^{2-}$		1.02×10^{-7}
Dihydrogen Phosphate Ion	$H_2PO_4^-$	$\rightleftharpoons$	$H^+ + HPO_4^{2-}$		6.23×10^{-8}
Boric	H_3BO_3	$\rightleftharpoons$	$H^+ + H_2BO_3^-$		7.3×10^{-10}
Ammonium Ion	NH_4^+	$\rightleftharpoons$	$H^+ + NH_4^+$		5.64×10^{-10}
Hydrogen Carbonate Ion	HCO_3^-	$\rightleftharpoons$	$H^+ + CO_3^{2-}$		5.61×10^{-11}
Monohydrogen Phosphate Ion	HPO_4^{2-}	$\rightleftharpoons$	$H^+ + PO_4^{3-}$		2.2×10^{-13}
Water	H_2O	$\rightleftharpoons$	$H^+ + OH^-$		1.0×10^{-14}

For space considerations, all reactions omit water as a reactant



19 3 Chemistry Answers Strengths Of Acids And Bases

The Princeton Review



19 3 Chemistry Answers Strengths Of Acids And Bases:

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