

Accelerated Curriculum for Mathematics Grade 7 TAKS

Student Edition



Accelerated Curriculum For Mathematics Grade 8 Taks

Sebastian Brünink



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$$\begin{aligned} & \text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3 \\ & 2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2 \\ & 2 \text{NaCl} + \text{F}_2 \rightarrow 2 \text{NaF} + \text{Cl}_2 \\ & 2 \text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O} \\ & \text{Pb} + \text{OH}^- + 2 \text{HCl} \rightarrow \text{H}_2\text{O} + \text{PbCl}_2 \\ & 2 \text{Al} + 3 \text{K}_2\text{SO}_4 \rightarrow 6 \text{KBr} + \text{Al}_2\text{SO}_4 \\ & 3 \text{CH}_4 + 2 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O} \\ & \text{C}_3\text{H}_8 + 5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O} \\ & \text{C}_8\text{H}_{18} + 25 \text{O}_2 \rightarrow 16 \text{CO}_2 + 18 \text{H}_2\text{O} \end{aligned}$$

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web may 28 2020 write a balanced molecular equation describing each of the following chemical reactions
 solid calcium carbonate is heated and decomposes to solid calcium oxide and carbon dioxide gas
 gaseous butane C_4H_{10} reacts with diatomic oxygen gas to yield gaseous carbon dioxide and water vapor

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 hen balancing

$$\begin{aligned} & 1 \text{Na} + 3 \text{PO}_4 + 3 \text{KOH} + 3 \text{NaOH} + 1 \text{K} + 3 \text{PO}_4 + 1 \text{MgF}_2 + 1 \text{Li} + 2 \text{CO}_3 + 1 \text{MgCO}_3 + 2 \text{LiF} + 1 \text{P}_4 + 3 \text{O}_2 + 2 \text{P}_2 + 3 \text{O}_3 + 2 \text{RbNO}_3 + 1 \text{BeF}_2 + 1 \text{Be} + \text{NO}_3 + 2 \text{RbF} \end{aligned}$$

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 $1\text{ N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF}$
 $\text{Cl}_2 + 4\text{H}_2\text{O} \rightarrow 2\text{H}_2 + 2\text{H}_2\text{O}_2$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $6\text{AlBr}_3 + 3\text{K}_2\text{SO}_4 \rightarrow 4\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{H}_2\text{O}$
 $4\text{CH}_4 + 2\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$
 $8\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

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 $1\text{O}_2 + 3\text{H}_2 \rightarrow 3\text{H}_2\text{O}$
 $2\text{O}_2 + 2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$
 $4\text{RbH} + 2\text{O} \rightarrow \text{RbOH} + 2\text{H}_2$
 $5\text{Cl}_2 + 2\text{KBr} \rightarrow \text{KCl} + \text{Br}_2$
 $6\text{N}_2 + 2\text{H}_2 \rightarrow 3\text{N}_2 + 7\text{N}_2\text{F}_3$
 $3\text{Ca} + \text{N}_2\text{O} \rightarrow \text{CaF}_2 + 8\text{CH}_4$
 $2\text{H}_2\text{O} + \text{CO} \rightarrow \text{H}_2 + 9\text{C}$
 $2\text{H}_2 + 5\text{OH}$

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 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF}$
 $\text{Cl}_2 + 2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $6\text{AlBr}_3 + 3\text{K}_2\text{SO}_4 \rightarrow 4\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{CH}_4$
 $2\text{O}_2 \rightarrow 1\text{CO}_2 + 2\text{H}_2\text{O}$

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 $2\text{FeBr}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 1\text{Fe}_2\text{SO}_4 + 3\text{H}_2\text{O} + 3\text{HBr}$
 $6\text{C}_4\text{H}_{10} + 3\text{O}_2 \rightarrow 2\text{C}_2\text{H}_4 + 2\text{O}_2$

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 $38\text{Fe} + 2\text{O}_3 \rightarrow 3\text{CO}_2 + 2\text{Fe}_3\text{CO}_2$
 $39\text{H}_2 + 3\text{PO}_4 \rightarrow 3\text{BaO} + 3\text{PO}_4$
 $4\text{H}_2 + 2\text{O} \rightarrow 4\text{O}_2 + 2\text{MgNH}_4$
 $4\text{PO}_4 + 2\text{P} \rightarrow 2\text{O}_7 + 2\text{NH}_3$
 $4\text{H}_2 + 2\text{O} \rightarrow 4\text{H}_2\text{O} + 2\text{Al}$
 $3\text{H}_2 + 2\text{SO}_4 \rightarrow 4\text{Al}_2\text{SO}_4 + 3\text{H}_2$
 $42\text{P}_4 + 10\text{H}_2\text{O} \rightarrow 3\text{HCl} + \text{POCl}_3$
 $3\text{HPO}_3 + 4\text{HNO}_2 + 2\text{H}_2\text{O} \rightarrow 3\text{O}_2 + 4\text{HNO}$

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