

Limiting Reactants Practice KEY

1. How many grams of aluminum chloride could be produced from 35.5 g of aluminum and 39.0 g of chlorine gas?



$$35.5 \text{ g Al} \times \frac{1 \text{ mol Al}}{26.982 \text{ g Al}} \times \frac{3 \text{ mol Cl}_2}{2 \text{ mol Al}} \times \frac{70.906 \text{ g Cl}_2}{1 \text{ mol Cl}_2} = 140. \text{ g Cl}_2$$

- ∞ This means 35.5 g Al would react with 140. g Cl₂
- ∞ Only 39.0 g Cl₂ were available, so it is the limiting reactant.
- ∞ Use the mass of Cl₂ in the reaction to determine the mass of the AlCl₃ formed.

$$39.0 \text{ g Cl}_2 = \text{ ? } \text{ g AlCl}_3$$

$$39.0 \text{ g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.906 \text{ g Cl}_2} \times \frac{2 \text{ mol AlCl}_3}{3 \text{ mol Cl}_2} \times \frac{133.341 \text{ g AlCl}_3}{1 \text{ mol AlCl}_3} =$$

Answer: 48.9 g AlCl₃

2. Identify the limiting reagent when 10.8 g water reacts with 4.7 g Na to produce NaOH and H₂.

To solve, set one reactant to the other and compare what is calculated to the actual amount in the reaction to determine the theoretical limiting reactant.

$$10.8 \text{ g H}_2\text{O} = \text{ ? } \text{ g Na}$$

Or

$$4.7 \text{ g Na} = \text{ ? } \text{ g H}_2\text{O}$$



$$10.8 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.015 \text{ g H}_2\text{O}} \times \frac{2 \text{ mol Na}}{2 \text{ mol H}_2\text{O}} \times \frac{22.990 \text{ g Na}}{1 \text{ mol Na}} = 13.8 \text{ g Na}$$

- ∞ This means that 10.8 g will react completely with 13.8 g Na, but only 4.7 g Na was available to this reaction. Therefore, Na is the limiting reactant in this problem.

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JS Bruner



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