

| Exam 1                            |                                                    |                                                                 | Exam 1 (cont)                    |                                 |                                                                                                           | Exam 1 (cont)                                                                                                                                                                                  |                                                                       |        | Exam 1 (cont)                                    |                                                                        |                                                                                 |
|-----------------------------------|----------------------------------------------------|-----------------------------------------------------------------|----------------------------------|---------------------------------|-----------------------------------------------------------------------------------------------------------|------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-----------------------------------------------------------------------|--------|--------------------------------------------------|------------------------------------------------------------------------|---------------------------------------------------------------------------------|
| Kelvin to Celsius                 | $K = C + 273$                                      |                                                                 | electrostatic energy             | $E_{el} = Q_1Q_2/d$             | Q1 and Q2: product of charges; d: distance between charges                                                | energy (hν) of a photon used to eject electrons from a metal surface via the photoelectric effect is equal to the sum of kinetic energy of the ejected electron (Ek) and the work function (W) | $h\nu = E_k + W$                                                      | Ek: hν | difference in energy between two quantum states  | $E = h\nu = -2.18 \times 10^{-18} J (1/n_f)^2 - 1/n_i)^2$              |                                                                                 |
| Fahrenheit to Celsius             | $F = 5F/9C (C) + 32F$                              |                                                                 |                                  |                                 |                                                                                                           | wavelength of emitted/absorbed light when an electron transitions from one quantum state to another                                                                                            | $1/\text{wavelength} = 1.097 \times 10^7 m^{-1} (1/n_f)^2 - 1/n_i)^2$ |        | energy of an electron with a given quantum state | $E_n = -2.18 \times 10^{-18} J (1/n^2)$                                |                                                                                 |
| density                           | $d = m/V$                                          | SI: kg/m <sup>3</sup> ; g/mL or g/cm <sup>3</sup> commonly used | joule                            | $1 J = 1 kg \times m^2/s^2$     | $1 J = 1 N \times m$                                                                                      |                                                                                                                                                                                                |                                                                       |        | wavelength of emitted/absorbed light             | $1/\text{wavelength} = 2.18 \times 10^{-18} J/hc (1/n_f)^2 - 1/n_i)^2$ |                                                                                 |
| moles to atoms and molecules      | 1 mole = $6.022 \times 10^{23}$ atoms or molecules |                                                                 | speed, wavelength, and frequency | $c = (\text{wavelength}) (\nu)$ | c: speed of light = $3.00 \times 10^8$ m/s; wavelength in meters; frequency (ν): in s <sup>-1</sup> or Hz |                                                                                                                                                                                                |                                                                       |        | de Broglie wavelength                            | $\text{wavelength} = h/mu$                                             | m: mass of particle in kg; u: velocity of the particle in s <sup>-1</sup> or Hz |
| moles to grams                    | 1 mole = atomic mass (g)                           | 1 mole = formula mass (g)                                       |                                  |                                 |                                                                                                           |                                                                                                                                                                                                |                                                                       |        |                                                  |                                                                        |                                                                                 |
| grams to atoms or molecules       | atomic mass (g) = $6.022 \times 10^{23}$ atoms     | formula mass (g) = $6.022 \times 10^{23}$ molecules             | energy of a photon               | $E = h\nu$                      | $h = 6.63 \times 10^{-34} J \times s$ ; ν: frequency in s <sup>-1</sup> or Hz                             |                                                                                                                                                                                                |                                                                       |        |                                                  |                                                                        |                                                                                 |
| avogadro's number                 | $6.022 \times 10^{23}$ moles                       |                                                                 |                                  |                                 |                                                                                                           |                                                                                                                                                                                                |                                                                       |        |                                                  |                                                                        |                                                                                 |
| kinetic energy of a moving object | $E_k = 1/2 mu^2$                                   | u: velocity                                                     |                                  |                                 |                                                                                                           |                                                                                                                                                                                                |                                                                       |        |                                                  |                                                                        |                                                                                 |



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