AP* CHEMISTRY 2013 SCORING GUIDELINES

Question 2

Answer the following questions involving the stoichiometry and thermodynamics of reactions containing aluminum species.

$$2 \text{ Al}_2\text{O}_3(l) + 3 \text{ C(s)} \rightarrow 4 \text{ Al}(l) + 3 \text{ CO}_2(g)$$

An electrolytic cell produces 235 g of Al(I) according to the equation above.

(a) Calculate the number of moles of electrons that must be transferred in the cell to produce the 235 g of Al(/).

235 g Al
$$\times \frac{1 \text{ mol Al}}{26.98 \text{ g Al}} = 8.71 \text{ mol Al}$$
Al³⁺ + 3 e⁻ \rightarrow Al, therefore, 3 mol e⁻ transferred per mol Al
8.71 mol Al $\times \frac{3 \text{ mol } e^-}{1 \text{ mol Al}} = 26.1 \text{ mol } e^-$

I point is earned for the number of moles of Al.

- 1 point is earned for correct stoichiometry and the number of moles of electrons.
- (b) A steady current of 152 amp was used during the process. Determine the amount of time, in seconds, that was needed to produce the Al(I).

charge = moles
$$e^- \times$$
 Faraday's constant
= 26.1 mol $e^- \times \frac{9.65 \times 10^4 \text{ C}}{1 \text{ mol } e^-} = 2.52 \times 10^6 \text{ C}$
 $I = \frac{q}{I}$
 $t = \frac{q}{I} = \frac{2.52 \times 10^6 \text{ C}}{152 \text{ C/s}} = 1.66 \times 10^4 \text{ s}$

I point is earned for the correct amount of charge transferred.

I point is earned for the correct time.

(c) Calculate the volume of CO₂(g), measured at 301 K and 0.952 atm, that is produced in the process.

$$mol\ CO_2 = 8.71\ mol\ A1 \times \frac{3\ mol\ CO_2}{4\ mol\ A1} = 6.53\ mol\ CO_2$$

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{(6.53\ mol)(0.0821\frac{L\ atm}{mol\ K})(301\ K)}{0.952\ atm} = 1.70 \times 10^2\ L\ CO_2$$

I point is earned for the number of moles of CO₂.

1 point is earned for the volume of CO₂.

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Marcel A. Müller

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