Section 9.2: Team Learning Worksheet

1. Consider the following unbalanced equation:

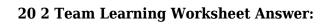
$$HCl + MnO_2 \rightarrow H_2O + MnCl_2 + Cl_2$$

You have 5.00 grams of manganese(IV) oxide:

- a. How many moles of manganese(IV) oxide do you have?
- b. How many moles of HCl do you need for a complete reaction?
- c. How many grams of HCl do you need for a complete reaction?
- d. How many grams of each product (H₂O, MnCl₂, and Cl₂) would be formed?
- e. Prove that mass has been conserved in this reaction.
- f. Is the number of atoms conserved? Support your answer.
- g. Is the number of molecules conserved? Support your answer.

20 2 Team Learning Worksheet Answer

S Joubert, P Laporte, A McFarland, M Oosten, H Vrooman



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