

PLANCHE 58

x 1

Veneridae

1. *Periglypta crispata*

(Deshayes, 1853)

2. *Pleammon isabellina*

(Philippi, 1849)

3. *Lioconcha ornata*

(Lamarck, 1817)

4. *Gafrarium pectinatum*

(Linné, 1758)

5. *Paphia textile*

(Gmelin, 1791)

6. *Samarangia quadrangularis*

(Adams & Reeve, 1850)

7. *Lioconcha tigrina*

(Lamarck, 1818)

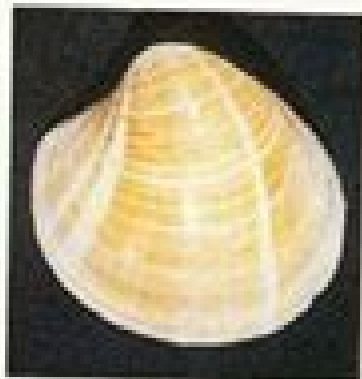
8. *Callista costata*

(Olivier, 1817)

Pitar affinis

(Gmelin, 1791) - 48 mm

Veneridae



1



2



3



4



5



6



7



8

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$$

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$$3\text{CO} + 7\text{H}_2\text{C} + 3\text{H}_2\text{O} \rightarrow 3\text{CH}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$
 $8.3 \text{ mole } \text{H}_2\text{O} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$
problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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 $\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g } \text{O}_2/\text{mole}} = 0.666 \text{ moles } \text{O}_2$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
b determine the theoretical yield of Al_2O_3
 $160.0 \text{ g } \text{Al} \times \frac{1 \text{ mol } \text{Al}}{27 \text{ g } \text{Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3/\text{mol} = 150.1 \text{ g } \text{Al}_2\text{O}_3$
c determine the percent yield
 $\frac{101.96 \text{ g } \text{Al}_2\text{O}_3}{150.1 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 67.9\%$

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web 7 Na + 2 C_2O_4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighed by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H_2 are left over 13

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